

Review

# Greenhouse Gas Sensors Fabricated with New Materials for Climatic Usage: A Review

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**Abstract:** With the increasing utilization of fossil fuels in today's technological world, the atmosphere's concentration of greenhouse gases is increasing and needs to be controlled. In order to achieve this goal, it is imperative to have sensors that can provide data on the greenhouse gases in the environment. The recent literature contains a few publications that detail the use of new methods and materials for sensing these gases. The first part of this review is focused on the possible effects of greenhouse gases in the atmosphere, and the second part surveys the developments of sensors for greenhouse gases with coverage on carbon nano-materials and composites directed towards sensing gases like CO<sub>2</sub>, CH<sub>4</sub>, and NO<sub>x</sub>. With carbon dioxide measurements, due consideration is given to the dissolved carbon dioxide gas in water (moisture). The density functional calculations project that Pd-doped single-walled carbon nanotubes are ideal for the development of NO<sub>x</sub> sensors. The current trend is to make sensors using 3D printing or inkjet printing in order to allow for the achievement of ppb levels of sensitivity that have not been realized before. This review is to elaborate on the need for the development of greenhouse gas sensors for climatic usage by using selected examples.

**Keywords:** greenhouse gases; carbon dioxide; methane; nitrogen oxides; fluorocarbons; optical methods; wireless sensors; resistive sensors; conducting polymer

## 1. Introduction

The entire world desires an atmosphere that provides clean air and has reduced levels of carbon dioxide, nitrous oxide, methane, ozone, and fluorocarbons, which contribute to global disturbances. These gases are called greenhouse effect gases. The atmosphere now faces two problems that contribute to the increased greenhouse effect gases. The first is the rapid deforestation caused by forest fires and land requirements for developmental purposes. The other is the use of fossil fuels for a variety of purposes. Fossil fuels generate gases that contribute to the greenhouse gas effect. The real concern is in the state of the global climate for the generations ahead. The United Nations (UN) has proclaimed several steps to mitigate greenhouse gas emissions. By this effort, it has reduced the burning of 64,000 kilotons of wood and reduced the release of 118,000 tons of CO<sub>2</sub> into the environment (<http://www.undp.org/content/undp/en/home/ourwork/climate-and-disaster-resilience/climate-change/mitigating-greenhouse-gas-emissions.html>). By encouraging the use of renewable sources of energy, several wind farms and solar panels have been developed to reduce greenhouse effect gases.

Among the greenhouse gases involved in environmental disturbances, carbon dioxide, nitrogen oxides, and methane are given utmost importance, as the concentrations of these gases change more than the others, thus requiring immediate attention. Hence, this review is devoted to these gases. The climatic conditions are influenced by greenhouse gases [1,2]. This study is not aimed at correlating

the gas levels to climate changes, although indirect relationships have been established, and hopefully this article will stimulate an interest in evolving such correlations.

Weather and climatic conditions are often considered important in our day-to-day life and, as a result, we hear/learn of these environmental conditions through the radio and TV. These channels often give us information such the temperature, wind speed, rain, snow, ice conditions, UV index, cloud cover, humidity, dew point, and barometric pressure, as shown in Figure 1. The weather channel seldom gives the CO<sub>2</sub>, CH<sub>4</sub>, N<sub>2</sub>O, or O<sub>3</sub> levels. This may be because of the non-availability of data and the methodology needed to be adapted to measure the concentrations of greenhouse gases. In 2012, a network of three tall tower stations was set up to obtain the long-term background levels of CO<sub>2</sub>, CH<sub>4</sub>, and NO<sub>x</sub> in the United Kingdom. The results obtained from the network showed that the greenhouse gases were increasing in mole fractions. The CO<sub>2</sub> and CH<sub>4</sub> concentrations showed nighttime maxima and daytime minima [3].



**Figure 1.** The weather conditions.

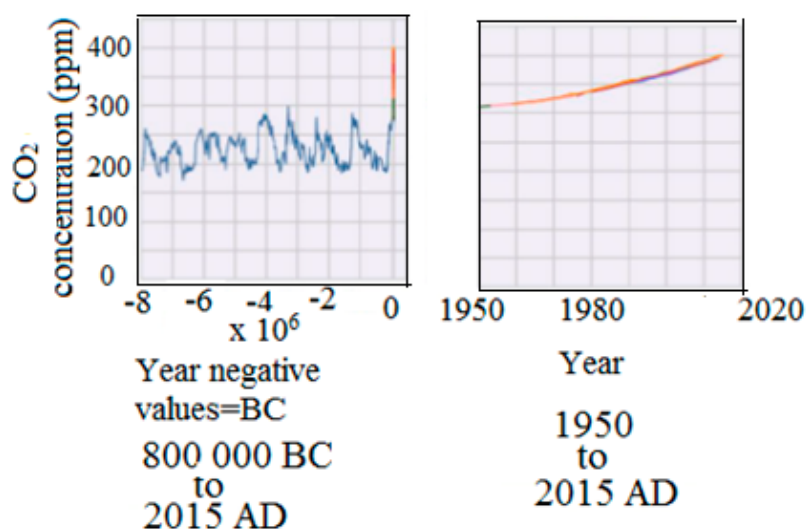
### 1.1. Status of Greenhouse Gases

The concentrations of greenhouse effect gases in the environment have been periodically monitored by several agencies, and in this section we review the concentrations of these gases in the environment with a view to developing sensors.

#### 1.1.1. Status of Carbon Dioxide Gas

The carbon dioxide level in the atmosphere is currently reported to be at about 400 ppm, compared to a low value of about 200 ppm in 800,000 BC (Figure 2). This reveals a rapid growth of the carbon dioxide levels in the atmosphere. This increase is attributed to fuel combustion, forest fires, volcanic eruption, and volatile organic compounds. The obvious choice to overcome this problem is to reduce the use of fossil fuels, increase the use of nonpolluting fuels, begin forest conservation efforts, and stop volcanoes from erupting [1,2]. The effectiveness of the measures to be taken requires sensors to monitor the carbon dioxide levels in the atmosphere. This has resulted in the evolution of gravimetric sensors. These sensors fall into the category of micro-to-nano electromechanical systems that are based on chemical, electrical conductivity, optical, magnetic, acoustic, capacitive,

or ultrasonic properties. The active material response to the exposed gas is revealed by the sensor's physicochemical interactions.

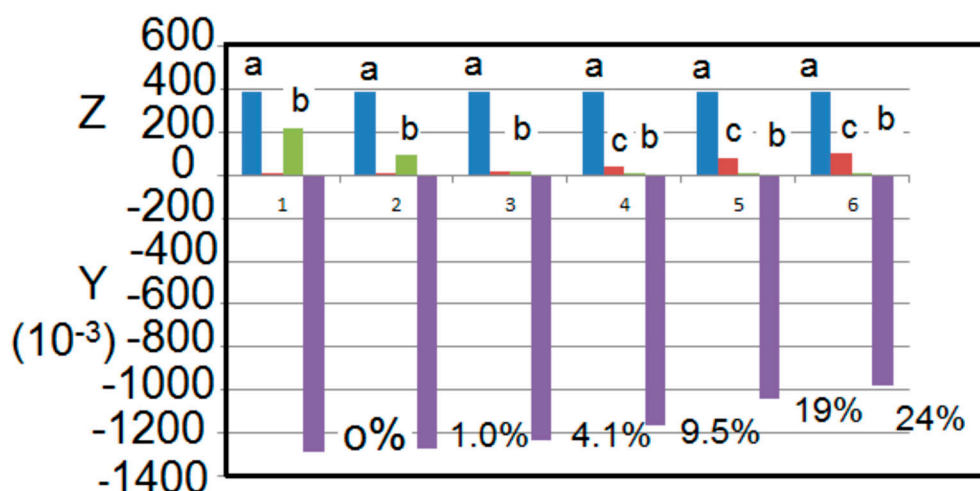


**Figure 2.** Global atmospheric concentrations of carbon dioxide over time. Taken from: <https://www.epa.gov/climate-indicators/climate-change-indicators-atmospheric-concentrations-greenhouse-gases>.

### 1.1.2. Status of Methane

Methane is another important gas contributing to the greenhouse gas effect. Its increase is greatly changing the climatic condition. It absorbs the sun's heat and warms the atmosphere. It is 84 times more potent than carbon dioxide in this activity. It is produced by the natural decomposition of rice paddies, marshes, the guts of animals, the rotting of rubbish and the distribution of fossil fuels like coal, oil, or gas. The use of alternatives to natural gas such as hydrogen in a variety of living conditions [4] could certainly reduce the contribution of this gas towards the greenhouse gas effect. Methane gas in the atmosphere has not been a serious concern, as it was increasing by 0.5 parts per billion per year until the year 2000. The current statistics show that it increased by 12.5 parts per billion in 2014. The methane mole fraction (ppb) is expected to reach about 1850 by 2018, and was approximately 1780 in the year 2000 ([https://www.washingtonpost.com/news/energy-environment/wp/2016/12/11/atmospheric-levels-of-methane-a-powerful-greenhouse-gas-are-spiking-scientists-report/?utm\\_term=.4a77e6472f2c](https://www.washingtonpost.com/news/energy-environment/wp/2016/12/11/atmospheric-levels-of-methane-a-powerful-greenhouse-gas-are-spiking-scientists-report/?utm_term=.4a77e6472f2c)). A recently published review [5] demonstrated that a sensor developed for carbon dioxide should consider its response to methane in the atmosphere. Figure 3 shows the simulated effect of methane interference in the measurement of carbon dioxide gas. The error in measurement of carbon dioxide concentration is indicated in the Figure for different concentration ratios of carbon dioxide to methane.

The effect of methane and carbon dioxide gases on the environment was recently analyzed by Charnay et al. [6] using a 3D climate carbon model, and their results showed a change in the global albedo from 0.40 to 0.23 depending on the relative proportions of carbon dioxide and methane levels. The global temperature was estimated by using the albedo value to increase from  $-11.5\text{ }^{\circ}\text{C}$  to  $65\text{ }^{\circ}\text{C}$ . This model predicts that a carbon dioxide level of 1 bar could produce hot climates at a low land fraction and cloud feedback. For the Earth to reach the high temperature requires carbon dioxide in the atmosphere to be 1 bar at 3.8 Ga.



**Figure 3.** Carbon dioxide sensor responses to different concentrations of methane. Y = Sensor response; Z = Concentration. Group 1 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 1.8 ppm; Group 2 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 4.25 ppm; Group 3 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 18 ppm; Group 4 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 40 ppm; Group 5 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 80 ppm; Group 6 a: CO<sub>2</sub> concentration, 390 ppm, b: Ratio of carbon dioxide/methane, c: methane concentration 100 ppm. Reproduced with permission from Reference [5]. Nonappearance of c in 1, 2 and 3 indicates negligible values.

### 1.1.3. Status of NO<sub>x</sub>

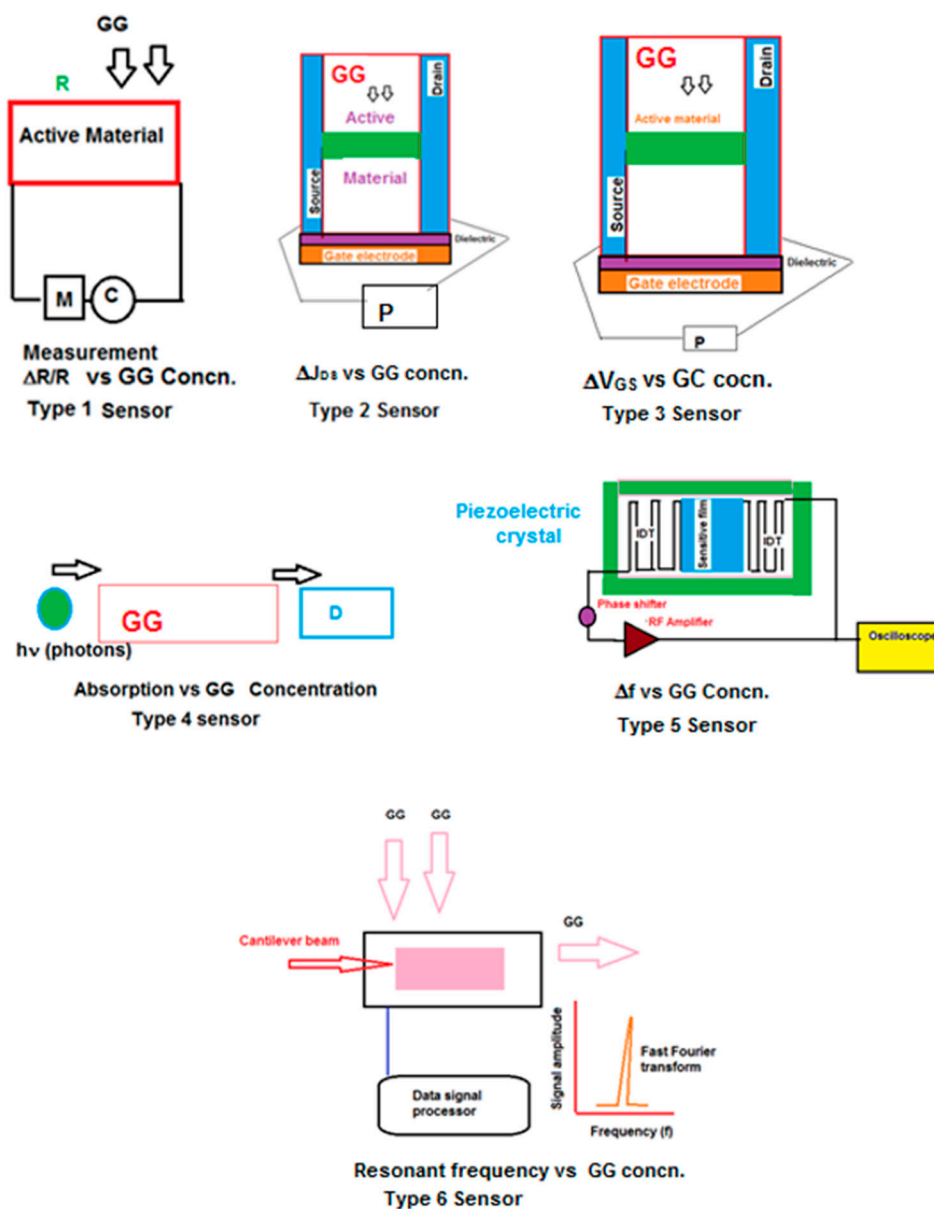
Nitrogen oxides have a pronounced influence in the environment. They are responsible for producing yellow-brown smog, often called photochemical smog. The sources for nitrogen oxides are automobile exhaust and fuel burning (including bio-diesel). It is to be controlled at 0.03 ppm over a one-year period for human health and environmental factors. Levels are currently estimated to be at 325 ppb in the environment (<https://www.epa.gov/climate-indicators/climate-change-indicators-atmospheric-concentrations-greenhouse-gases>).

### 1.1.4. Status of Fluorocarbons

The fluorocarbons are innumerable, and they can be divided into two classes: one class acting on the ozone layer that reduces the ultraviolet rays reaching the Earth, and another class having environmental effects. The first-class compounds are methyl chloroform, halon-1211, CFC-12, HCFC-22, and HCFC-141b. The other class includes sulfur hexafluoride, HFC-23, HFC-134a, PFC-14, PFC-16, and nitrogen trifluoride. The ozone-depleting fluorocarbons, except for methyl chloroform, are either increasing or reaching a constant value. Their levels range from about 5 to 400 parts per thousand (ppt) in the atmosphere (<https://www.epa.gov/climate-indicators/climate-change-indicators-atmospheric-concentrations-greenhouse-gases>). Since fluorocarbons are not at the concentration levels of CO<sub>2</sub> or CH<sub>4</sub>, they will not be considered here.

## 1.2. Sensors for Greenhouse Gases

Since the importance of the influence of greenhouse gases on the environment was understood, several different sensors have been developed and are well-reviewed in the literature [7–12]. Figure 4 describes the six different types of sensors used in the measurement of greenhouse gases.



**Figure 4.** Types of sensors developed for greenhouse gas measurements. M: digital multimeter; C: computer with software for monitoring the measuring multimeter; P: potentiostat (monitors the field effect transistor parameters); D: detector (e.g., photomultiplier or bolometer); GG: greenhouse gas; IDT: inter digitized electrode.

Type 1 sensors generally have a nonmetallic substrate on which the active material is deposited and kept in either ambient or selected experimental conditions (e.g., inert atmosphere) in the measuring chamber. The change in resistance of the active material is measured as a function of concentration of the greenhouse gas. The successive measurement requires flushing the measuring chamber. The type 2 sensor is a field effect transistor having active material placed between the source and drain. The source, drain, and gate electrodes are connected to a potentiostat. The drain current is measured as a function of greenhouse gas concentration. The type 3 sensor is a modification of the type 2 sensor, in that the voltage is plotted as a function of greenhouse gas concentration. The type 4 sensor depicts the general scheme for electromagnetic interaction with the greenhouse gas. The photons used in these measurements are generally in the energy range of  $10^4$ – $10^{-2}$  kcal. Type 5 sensors utilize a piezo electric crystal carrying interdigitized electrodes. The active material is placed between digitized electrodes. The radio frequency ( $R_f$ ) shift is measured with different greenhouse gas concentrations. This type of

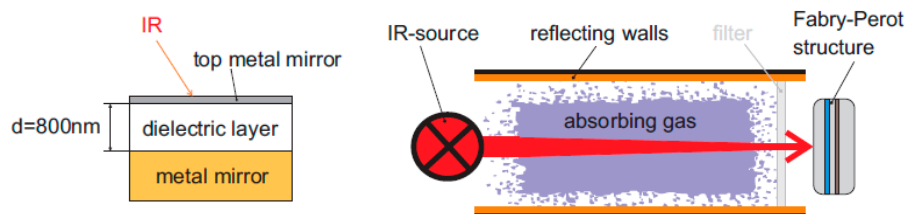
sensor is also called a SAW (surface acoustic wave) sensor. The type 6 sensor is a micromechanical sensor using a cantilever beam which has a coating of an active material. The resonant frequency of the cantilever beam is measured as a function of the greenhouse gas concentration. This review contains examples of all six different sensor types.

Conventional instrumental analysis such as optical spectroscopy, Fourier transform infrared spectroscopy (FTIR), semiconducting devices, mass spectroscopy, and Raman spectroscopy were used to analyze the greenhouse gases. These techniques will continue to be used, as they provide the advantages of selectivity and sensitivity. With the discovery of new nanomaterials, several less-expensive methods such as using resistance measurement (Type 1 sensor), field effect transistor (Types 2 and 3 sensors), optoacoustic (Type 4 sensor), wireless measurements (Type 5 sensor; surface acoustic wave (SAW)) and micromechanical measurements (Type 6) have recently been developed. These new methods and new research oriented towards new materials have provided speed and accuracy in the measurement of greenhouse gases.

## 2. Carbon Dioxide Sensors

Carbon dioxide gas is safe for humans up to 5000 ppm, and is dangerous when it reaches concentrations of 40,000 ppm. A short-term exposure of 30,000 ppm is bearable (<http://www.cdc.gov/niosh/idlh/intridl4.html>) [13,14]. The level of carbon dioxide in the atmosphere is changing every year. The present level of carbon dioxide in the atmosphere is over 400 ppm. The presence of it in the atmosphere affects the albedo value of the Earth in reflecting the solar radiation and is currently estimated at 0.39. While the planet Mercury has an albedo value of 0.1 (receives the maximum amount of sunlight), the planet Venus has a value 0.84 [4]. The temperature of the Venus is estimated at 462 °C, with carbon dioxide levels in the atmosphere amounting to 96.5%. Based on these facts, the need for limiting carbon dioxide in the Earth's atmosphere is of the utmost importance. Hence, a good and reliable sensor for carbon dioxide is required to monitor the atmospheric carbon dioxide. A variety of sensors have been developed based on optical absorption, change in semiconducting property, electrical resistance, amperometry, and field effect transistors. The semiconducting materials used in these sensors operate at high temperatures (>200 °C). The optical detectors use either fiber optics methods with sophisticated instrumentation or conventional infrared detectors. The electrical methods have been successfully used for the detection of carbon dioxide in biomedical applications.

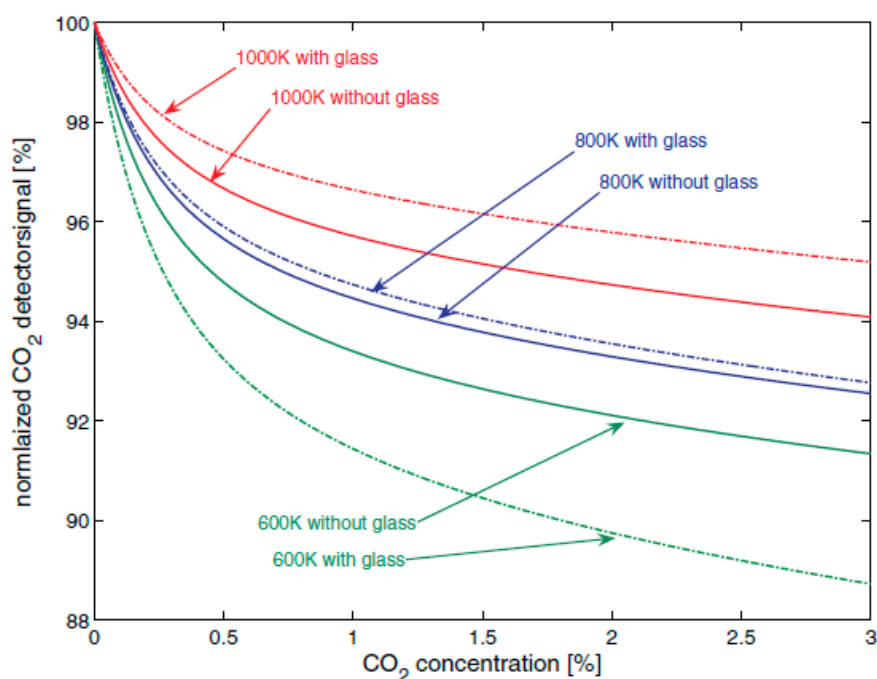
The main thrust in this decade has been to find new materials that enable easy detection and determination. Gravimetric sensors have been of great interest in this category. The infrared sensors are developed using a combination of a wavelength filter and a detector. This technology can reach a detection limit of 10 ppm of CO<sub>2</sub>, with an upper limit of 10,000 ppm [15–18]. Mayrwögera et al. [15] proposed a Fabry–Perot-based bolometer using a glass plate as a simple infrared filter for carbon dioxide determination (Type 4 sensor). Figure 5 shows the analyte that is mixed with nitrogen for analysis.



**Figure 5.** The dielectric layer used for the measurement and the methodology. Ge is used as a dielectric layer. Reproduced with permission from Reference [15]. IR: infrared radiation.

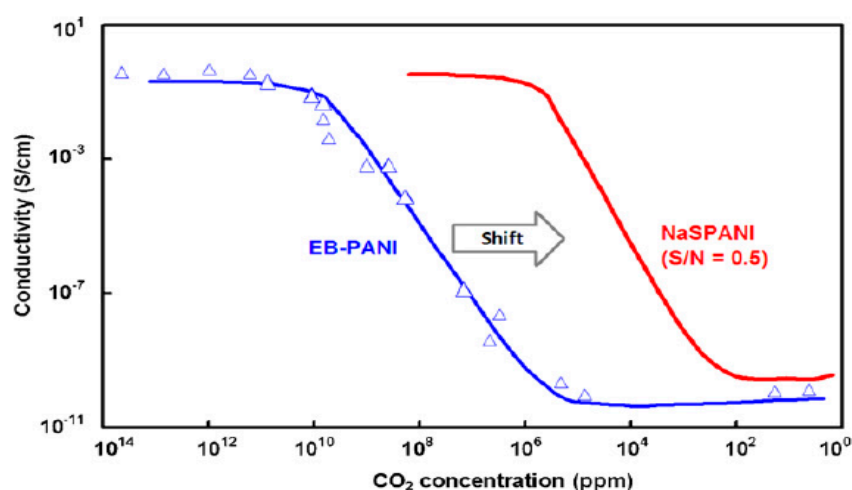
The interference coming from water vapor in the measurement of carbon dioxide concentration was removed by using a glass filter, as shown in Figure 6.





**Figure 6.** Carbon dioxide determination using glass filters. Reproduced with permission from Reference [15].

Several conducting polymer-based resistive sensors have been developed for the detection of carbon dioxide. Doping polyaniline (PANI) has been shown to play a role in the detection of  $\text{CO}_2$  [18–20]. The working range of the  $\text{CO}_2$  sensor has been reported as 102–104 ppm. Figure 7 shows the conductivity change with the concentration of  $\text{CO}_2$ . The conductivity decreases when carbon dioxide is adsorbed on the sensor material (PANI; Type 1 sensor). The sensor performance has been shown to depend on whether it has an emeraldine base or sulfonated polymer as its active material. Both of them respond to carbon dioxide, but the emeraldine base's response to lower ppm levels has been reported to be negligible.



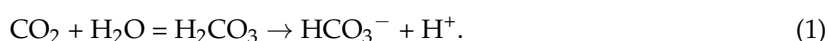
**Figure 7.** Conductivity of sensor with concentration of carbon dioxide. Reproduced with permission from reference [19]. EB-PANI: emeraldine base polyaniline. NaSPANI: sodium sulfonated PANI.

The carbon dioxide concentrations in the atmosphere and the corresponding pH are controlled by the humidity in the atmosphere. The dissolved carbon dioxide is in equilibrium with other species [19,20], and Table 1 shows the calculated concentrations of the equilibrating species.

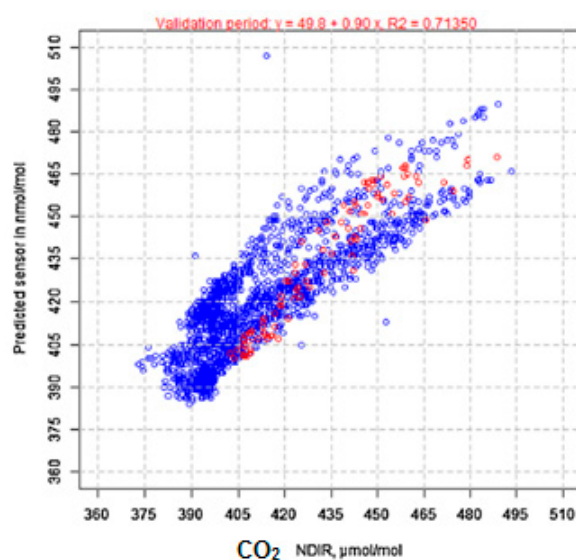
**Table 1.** The CO<sub>2</sub> concentrations in the atmosphere, pH value, and equilibrating species concentrations.

CO <sub>2</sub> (Gas, ppm)	pH	Concentration [CO <sub>2</sub> ]	Concentration [H <sub>2</sub> CO <sub>3</sub> ]	Concentration [HCO <sub>3</sub> <sup>-</sup> ]	Concentration [CO <sub>3</sub> <sup>2-</sup> ]
1.0 × 10 <sup>-1</sup>	6.94	3.36 nM	5.71 pM	5.90 nM	1.90 pM
1.0 × 10 <sup>0</sup>	6.81	33.6 nM	57.1 pM	91.6 nM	33.0 pM
1.0 × 10 <sup>1</sup>	6.42	0.33 μM	5.71 nM	0.378 μM	55.7 pM
1.0 × 10 <sup>2</sup>	5.92	3.36 μM	5.71 nM	0.378 μM	56.0 pM
1.0 × 10 <sup>3</sup>	5.42	33.6 μM	57.1 nM	3.78 μM	56.1 pM
1.0 × 10 <sup>4</sup>	4.92	0.336 mM	0.571 μM	0.119 μM	56.1 pM
1.0 × 10 <sup>5</sup>	4.42	3.36 mM	5.71 μM	0.378 μM	56.1 pM
1.0 × 10 <sup>6</sup>	3.92	33.6 mM	57.1 μM	12.0 mM	56.1 pM

The true concentration of CO<sub>2</sub> in the atmosphere can be evaluated by taking the concentrations in Table 1 into consideration, as there is equilibrium between CO<sub>2</sub> and the protonated species as



An infrared fiber optic optical spark plug sensor was used for measuring CO<sub>2</sub> and water. As both the molecules are infrared-active with strong overlap, the spark plug was kept close to the electrodes for the in-cylinder measurement of CO<sub>2</sub> and gaseous water (Type 4 sensor). A tungsten halide lamp with two infrared detectors having different optical band pass filter was used. The test was carried out using a spark-ignited engine [21]. Air quality monitoring sensors using a cluster of metal oxide (MO<sub>X</sub>; MiCS-5525) [22] or electrochemical sensors were used for the analysis of a mixture of nitrogen monoxide and carbon dioxide. The measurements were used to assess the practicality of the carbon dioxide sensor for a data quality objective with 25% uncertainty [23]. Table 2 shows the sensors used for data acquisition. Figure 8 shows the carbon dioxide levels obtained from the sensor with linear regression analysis.



**Figure 8.** Red points indicate measured data, and blue indicate simulated data. Reproduced with permission from Reference [23]. NDIR: non-dispersive infrared technology.



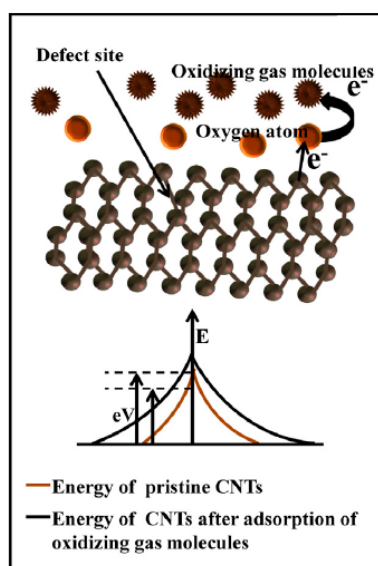
**Table 2.** Commercialized carbon dioxide sensors for field experiments.

Marketed by	Model	Concentration Range	Sensitivity	Response Time	Resolution	Comments
ELT Sensors	S-100H	0–5 mmol/mol	$\frac{1}{V/\text{mmol/mol}}$	60 s	70 nmol/mol	Uses 16 bits ADC with DAQ range of 5 V
Edinburgh sensors	Gas card NG	0–1 mmol/mol	$\frac{1}{V/0.1 \text{ mmol/mol}}$	10 s	60 nmol/mol	Uses 14 bits ADC with DAQ range of 10 V

Resistive sensor. (Type 1 sensor)  $R_s$ : resistance of the sensor in ambient air;  $R_o$ : is the resistance when not exposed to air. ADC: analog-to-digital converter; DAQ: data acquisition.

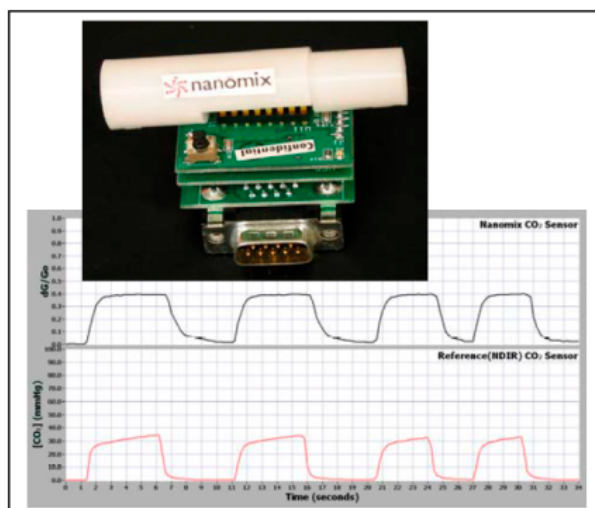
It was shown that the predicted values were lower compared to calibration. When an electrochemical sensor was used for the measurement of  $\text{CO}_2$ , the interference of ozone was encountered and was removed by linear/multi linear regression [23,24].

Carbon nanotube (CNT)-based sensors have been developed for carbon dioxide gas utilizing the principle of physicochemical adsorption of the gas by the carbon nanotubes. Two types of sensors have been developed based on this principle. One type is based on resistance change upon adsorption of the gas on the active material, and the other is dependent on the adsorption of the gas on the active material having effects on transistor properties such as the voltage or current in the field effect transistor [13] (Types 2 and 3 sensors). The adsorption of carbon dioxide gas produces an increase in resistance that is attributed to an increase in the energy barrier for electron movement, as shown in Figure 9.



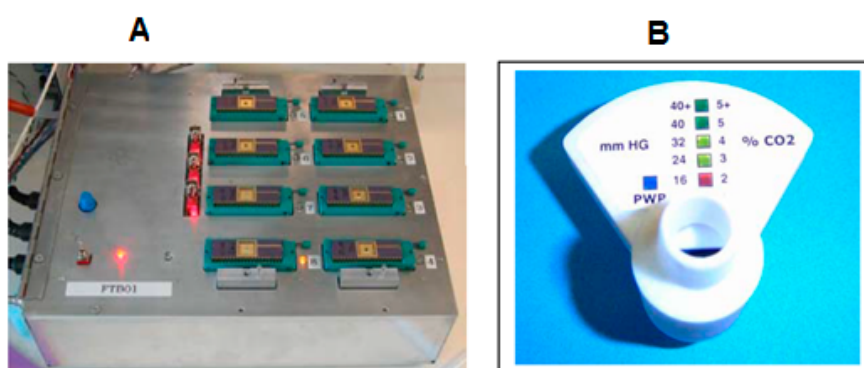
**Figure 9.** Sensing mechanism for carbon dioxide molecule. Reproduced with permission from Reference [13]. CNT: carbon nanotube.

A number of reports directed towards measuring  $\text{CO}_2$  using a chemoresistive method with CNTs have been reported in the literature [13,15]. Trans et al. [25] have fabricated a field effect transistor using carbon nanotubes (NTFET) whose sensitivity was examined by Star et al. [26–28]. A prototype sensor chip was packaged for measuring the carbon dioxide level using a computer. Figure 10 shows the response of the sensor for carbon dioxide in breath analysis. The reproducibility in the response to successive injections of the gas was established by these measurements. The NTFET board containing the field effect transistor [29,30] is shown in Figure 11. The reproducibility of the pulses shows the feasibility of its usage for successive measurements.



**Figure 10.** Field effect transistor using carbon nanotubes (NTFET) sensor fabricated for carbon dioxide. Reproduced with permission from Reference [28].

The sensor monitors the %CO<sub>2</sub> that can be detected by the NTFET, which can be used conveniently in several locations, such as in hospitals and on-site by paramedics. Being a noninvasive and disposable device, it can be used in the monitoring of greenhouse gas for about 6 h. The performance of the NTFET in measuring the concentration of CO<sub>2</sub> in the presence of moisture has also been carried out. It had a tolerance up to 80% RH (relative humidity). Sensors made with CNTs for gases generally have the following advantages: (a) room-temperature operation; (b) facile property adjustment; (c) high sensitivity and response time; (d) easy device fabrication; (e) low selectivity; and (f) long time instability [10,31–39].

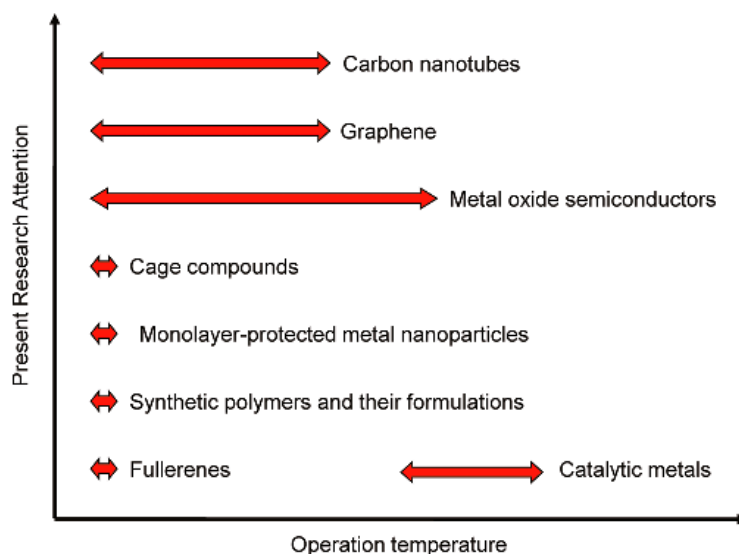


**Figure 11.** (A) Functional test board (FTB) loaded with eight packaged die containing functionalized NTFET devices. (B) Mockup capnography device. Reproduced with permission from reference [28].

The low selectivity and long time instability have been an active area of investigation in the past decade, and several new composites of carbon nanotubes have been recommended for gas detection. Chen et al. [19] have reported that cleaning the carbon nanotubes with ultraviolet light resulted in a dramatically enhanced performance. The mechanism by which ultra violet light enhances the performance has not been explored in this work. Presumably it makes the surface active by the oxidation of amorphous carbon or defect centers in the nanotubes. A quantum mechanical investigation of carbon dioxide adsorption on S-functionalized boron nitride and aluminum nitride (AlN) nanotubes has shown it to be exothermic, opening the prospects of developing a thermal sensor for carbon dioxide [40]. The adsorption of carbon dioxide on AlN nanocages and nanosheets has also been investigated using density functional calculations [41,42], with emphasis on chirality.

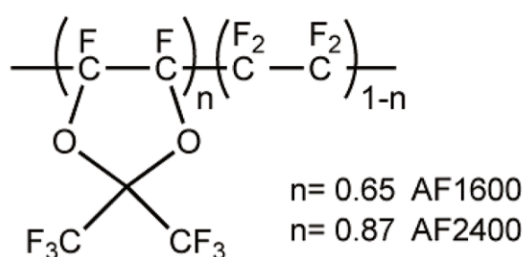
Amine-functionalized TiO<sub>2</sub> has been used to capture the carbon dioxide from flue gases, and is amenable for the development of a gas sensor. The presence of moisture has been demonstrated to have a significant effect on adsorption/desorption processes [43].

A nano thin film functionalized chemiresistor sensor operating at room temperature (Type 1 sensor) has been reported for the detection of carbon dioxide in the range of 50–500 ppm [44]. The sensor had negligible interference from ammonia, carbon monoxide, methane, and nitrogen dioxide. The sensing of carbon dioxide using a wireless network [13,45–47] (Type 5 sensor) uses analyte-induced changes such as in mass, elasticity, or complex permittivity [48]. The diverse temperature requirements in this class of carbon network are shown in Figure 12.

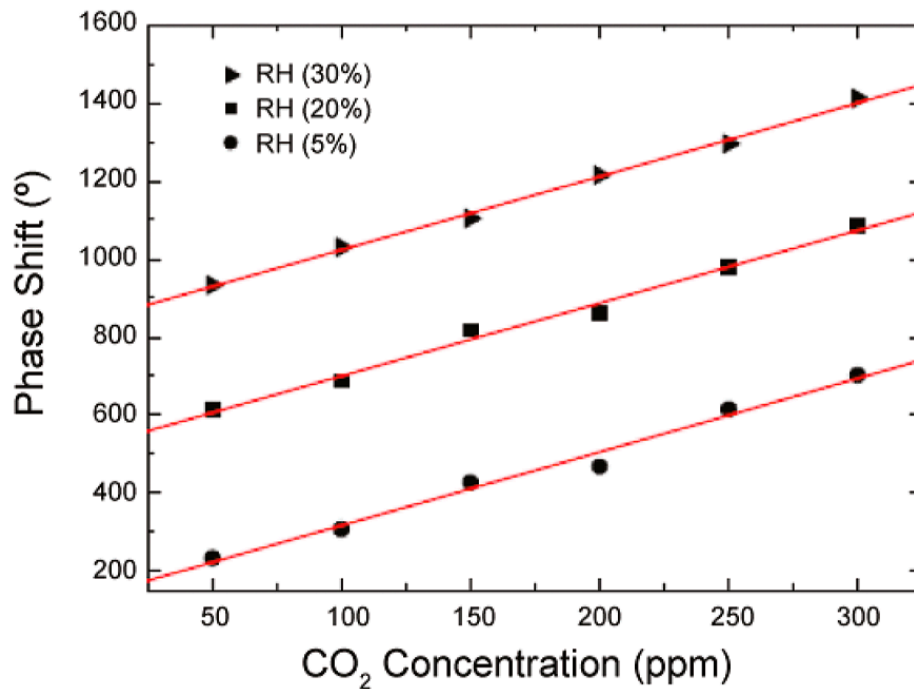


**Figure 12.** The temperature of sensor operation in relative scale. Reproduced with permission from Reference [48].

A SAW (surface acoustic wave) sensor using random copolymer Teflon AF2400 (Figure 13) prepared from tetrafluoroethylene and 2,2-bis(trifluoromethyl)-4,5-difluoro-1,3-dioxole showed variations in carbon dioxide detection depending on the amount of water present. The phase shift changes with carbon dioxide level are shown in Figure 14, showing a phase shift of about 1.14°/ppm of CO<sub>2</sub>.

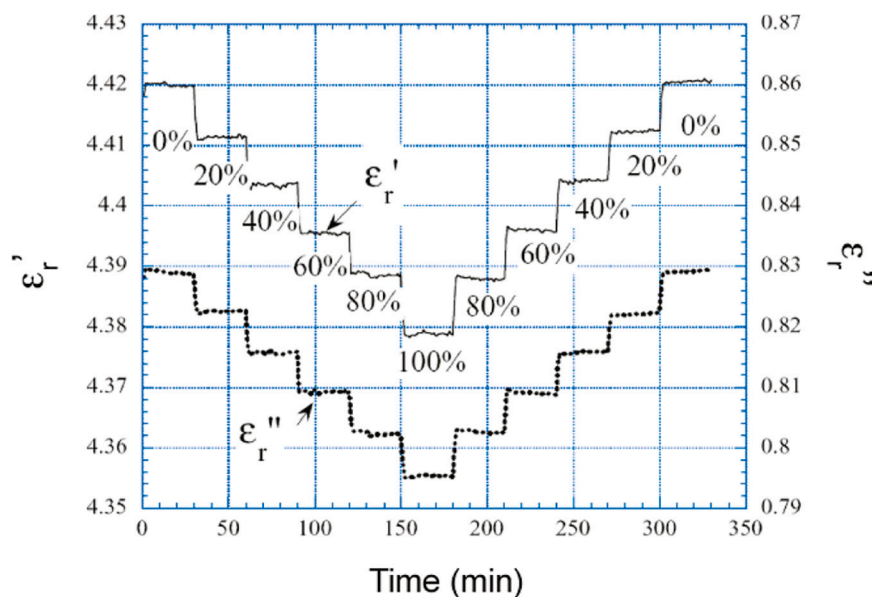


**Figure 13.** The structure of Teflon AF2400.



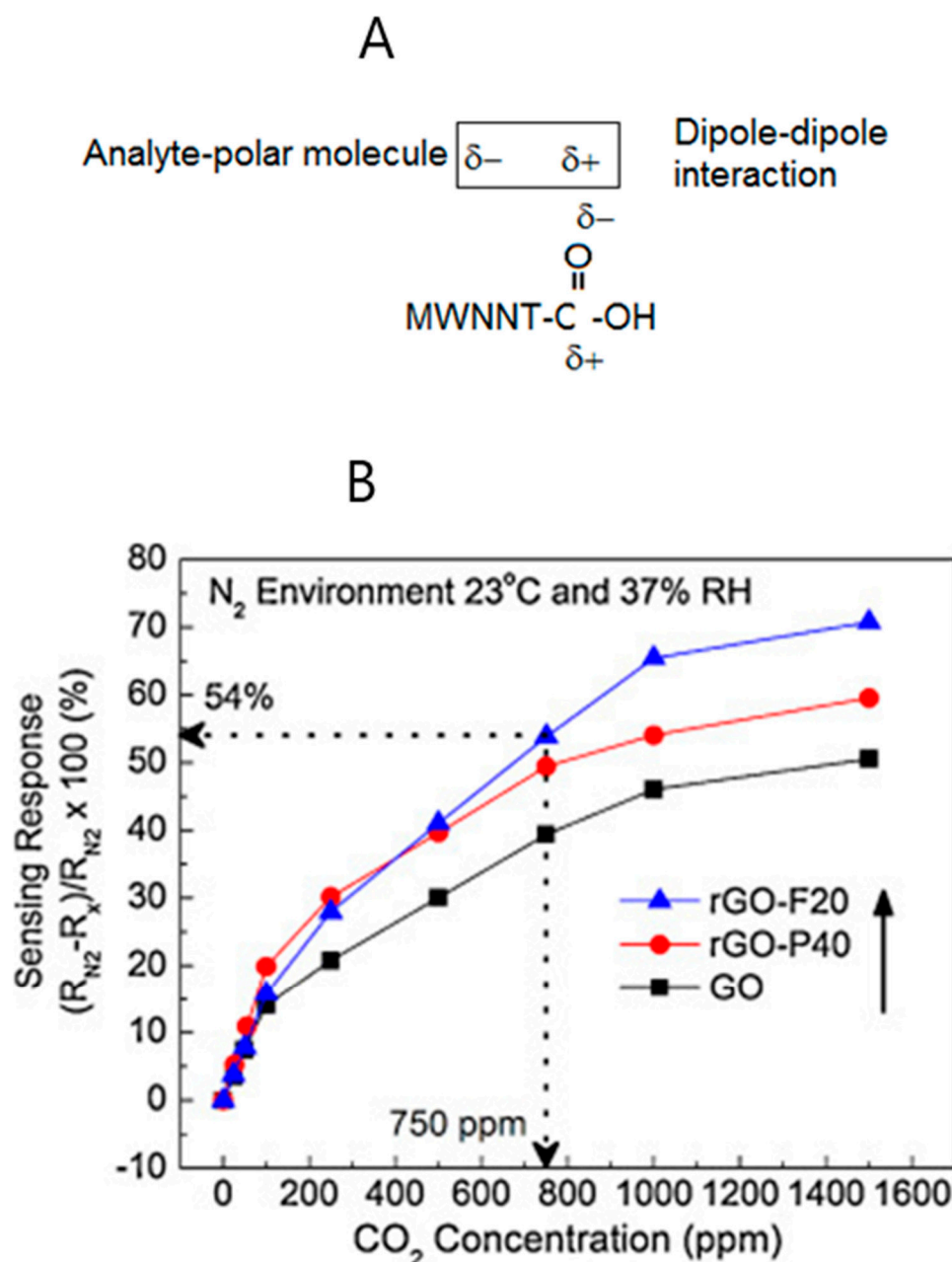
**Figure 14.** Phase shift observed as a function of carbon dioxide concentration at different water levels (RH: relative humidity). Reproduced with permission from Reference [48].

The permittivity and conductivity of multiwalled carbon nanotubes has been used for the detection of carbon dioxide [48]. The incorporation of a SiO<sub>2</sub> matrix along with the incorporation of a wireless inductor-capacitor resonator in the sensor showed a decrease in the effective permittivity (see Figure 15). The analyte induced changes in complex permittivity ( $\epsilon_r' - j\epsilon_r''$ ), where  $\epsilon_r'$  and  $\epsilon_r''$  are the real and imaginary parts of the complex permittivity are measured in this approach. The imaginary permittivity of the sensing material is directly proportional to its conductivity. A hysteresis-free operation of the sensor is remarkable and advantageous for fast measurements.



**Figure 15.** Carbon dioxide sensor using a wireless transducer constructed with multiwalled carbon nanotube (MWCNT)-SiO<sub>2</sub> film. Reproduced with permission from Reference [10].

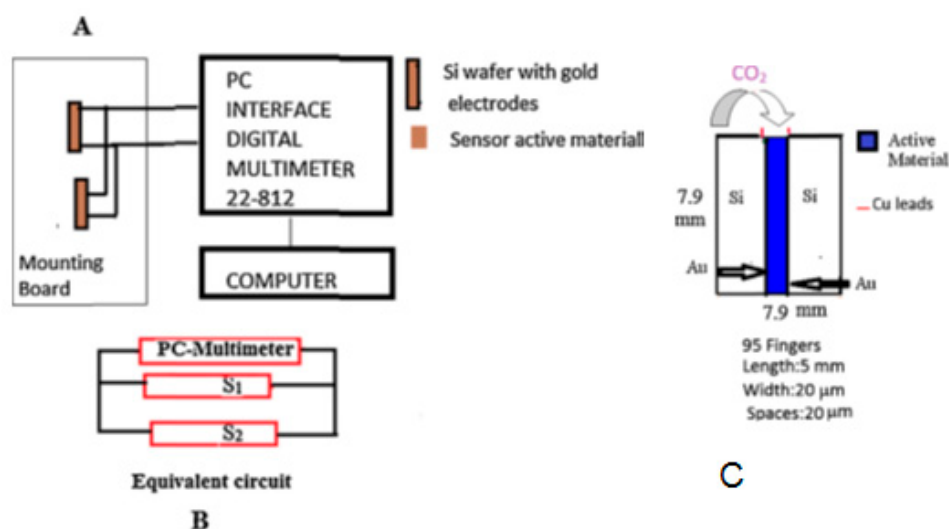
The sensor operates through a dipole–dipole interaction mechanism as shown below (Figure 16A). The functionalized multiwalled carbon nanotubes (MWCNTs) have a partial negative charge to which the analyte is bridged through its partial positive charge end, as shown in the above illustration [10]. The optical sensing of carbon dioxide has been reported using different types of membranes [38]. As an offshoot of carbon nanotubes, the discovery of graphene by Novoselov and Geim in 2004 has opened up new sensors based on carbon [39]. In one approach, 3  $\mu\text{L}$  of a 1 mg/mL graphene oxide solution was spin-coated on silicon fingers of 3  $\mu\text{m}$  width. The graphene resistive sensor's response in the concentration range of 0–1500 ppm was measured [49–52]. Figure 16B provides the resistance response with carbon dioxide concentration at ppm levels.



**Figure 16.** (A) Dipole-dipole interaction mechanism. (B) Graphene resistive sensor. Reproduced with permission from reference [49]. rGO: reduced graphene oxide; GO: graphene oxide.

The initial resistance measurement of the sensor was carried out in nitrogen atmosphere (RN<sub>2</sub>). The change in resistance [46] of the sensor when carbon dioxide gas was injected (Rx) was used in the construction of the graph in Figure 16. Recently, a miniature resistive carbon dioxide sensor has been reported that operates in the concentration range of 50–50,000 ppm. A photoacoustic spectroscopic method has been developed for the simultaneous determination of carbon dioxide and methane gases [52] with high precision and a large dynamic range.

A new silicon substrate micro sensor has been developed [5] using a composite made of carbon nanotube and Baytron-P that senses carbon dioxide at 22 °C. The sensor was constructed with a Si chip by depositing the composite between two gold electrodes. Two identical Si chips were connected in a parallel configuration (Figure 17) to reduce the initial resistance of the sensor. The resistance of the sensor decreased proportionally upon exposure to the concentration of carbon dioxide.



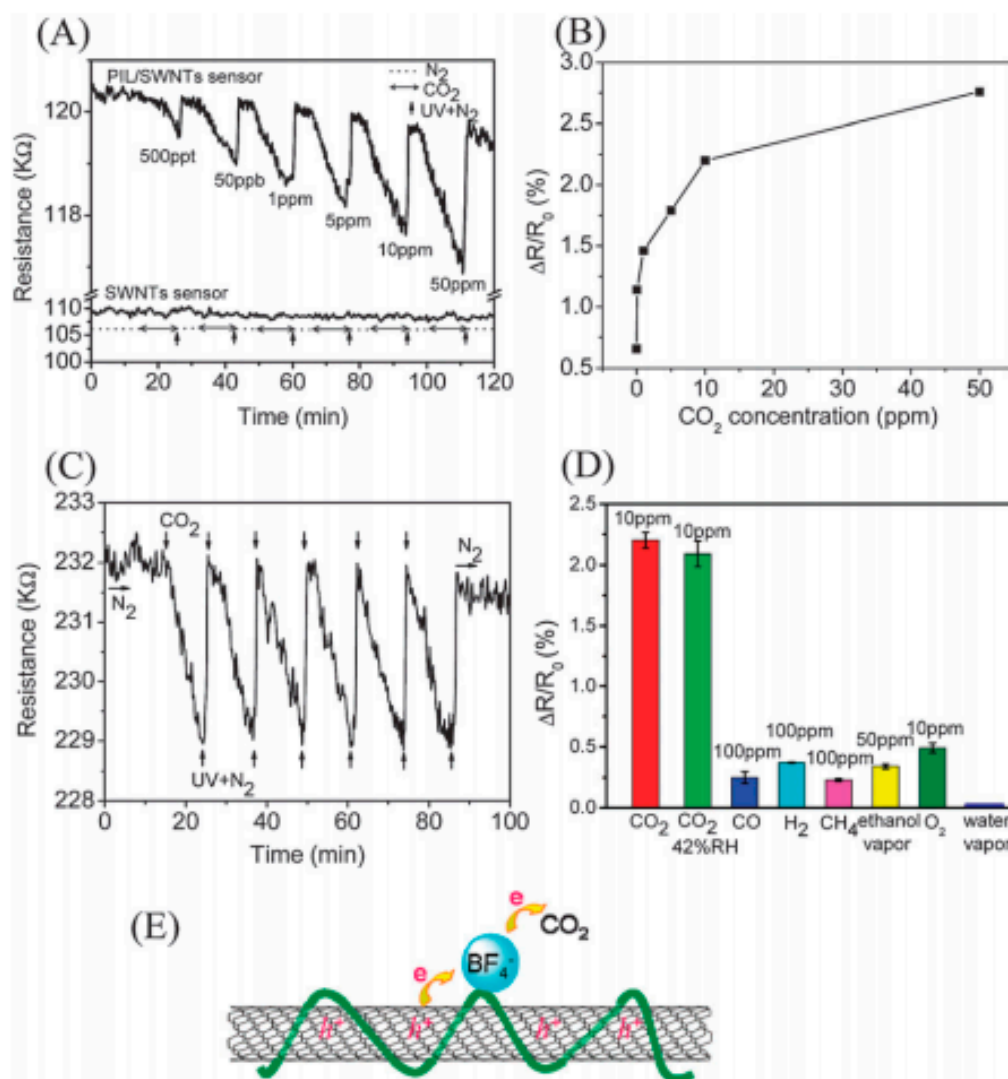
**Figure 17.** Twin sensors in parallel configuration. (A) The arrangement of the sensors and the measurement details. (B) The equivalent circuit of the arrangement of the sensors (S1 and S2). (C) Sensor dimensions. Reproduced with permission from Reference [5].

The sensor showed a semiconducting behavior with a negative temperature coefficient (Type 1 sensor). The response time of the sensor was about 40 s. Fourier transform infrared spectroscopy showed peaks for the nanocomposite at 1056  $\text{cm}^{-1}$ , 1195  $\text{cm}^{-1}$ , 1296  $\text{cm}^{-1}$ , 1635  $\text{cm}^{-1}$ , 2083  $\text{cm}^{-1}$ , 2345  $\text{cm}^{-1}$ , and 3278  $\text{cm}^{-1}$ . The carbon dioxide adsorption on the composite resulted in the polystyrene sulfonate absorption band shifting from 1195  $\text{cm}^{-1}$  to 1176  $\text{cm}^{-1}$ , suggesting a phase separation occurring in the nanocomposite that resulted in the increased conductivity.

Metal oxide semiconductor (MOS) films, nanowires, nanocage, powder, and microspheres have been investigated [53–69] for carbon dioxide sensing. The sensors developed using this approach operate at a temperature range of 200–700 °C and can detect concentrations in the range of 100–10,000 ppm with a response time falling in the range of 3 s to 9.5 h. The recovery times are also in the range 4–700 s. The performance of an MOS sensor depends on its morphology and composition. The results obtained here have shown that the grain size significantly influences the performance sensitivity (Types 2 and 3 sensors).

Poly (ionic liquid)-wrapped single-walled carbon nanotubes have been found to be sensitive for carbon dioxide detection at low concentrations of CO<sub>2</sub> (500 ppt) [67–72]. The chemiresistive dynamic response of the carbon nanotubes is shown in Figure 18 (Type 1 sensor), where the mechanism for sensing is based on the interaction between BF<sub>4</sub><sup>-</sup> anion and CO<sub>2</sub>. The charge transfer interaction between BF<sub>4</sub><sup>-</sup> and CO<sub>2</sub> is depicted in Figure 18E.



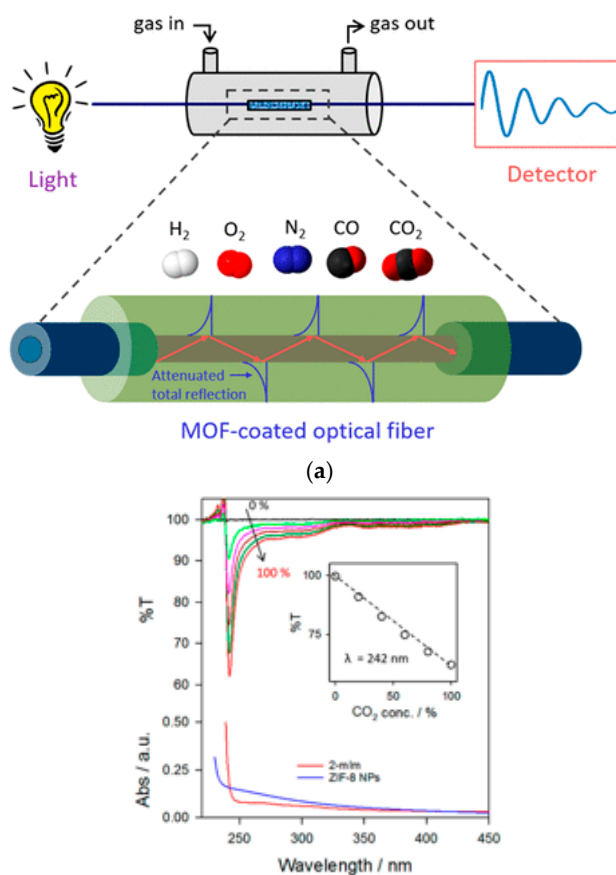


**Figure 18.** (A) shows the response of pristine single walled carbon nanotubes (B) Change in resistance with respect to  $CO_2$  concentrations. (C) Successive injections of 1 ppm  $CO_2$  to the sensor. (D) Response of the sensor to:  $CO_2$  (10 ppm);  $CO_2$  (10 ppm) + relative humidity (42%); 100 ppm  $H_2$ ; 100 ppm  $CH_4$ ; 50 ppm ethanol; 10 ppm  $O_2$  and relative humidity (42%). (E) Sensor mechanism. Reproduced with permission from [67].

An impressive type 4 sensor for  $CO_2$  has been reported using optical fiber coated with a metal oxide-zeolitic imidazole framework (ZIF)-8 MOF [69].

A high sensitivity for  $CO_2$  with negligible response for gases such as  $H_2$ ,  $N_2$ ,  $O_2$ , and  $CO$  has been reported. The percentage transmission of 242 nm radiation has been monitored as a function of  $CO_2$  concentration, as shown in Figure 19b. The linear change in the transmission of radiation with concentration with very little interference from other gases provides reliability for usage in the field.

Table 3 gives a list of sensor materials and their responses.



**Figure 19.** (a) CO<sub>2</sub> sensor with metal oxide-zeolitic imidazole framework on the optical fiber. Reproduced with permission from Reference [69]. (b) Performance characteristics of CO<sub>2</sub> sensor giving percentage transmission and absorbance features. MOF: Metal organic framework. Reproduced with permission from Reference [69].

**Table 3.** Carbon dioxide sensing materials.

Material	Detection Method	Operating Temperature (°C)	Detection Concentration	Response Time	Reference
Carbon nanotubes	Resistance		0–2000 ppm	3 s	[13]
Gascard-NG	Current–Voltage		0–1000 μmol/mol	10 s	[23]
Germanium	IR sensor	30–70	10 ppm		[15,18]
PANI	Resistance	Room temperature	10 <sup>2</sup> –10 <sup>5</sup> ppm		[18,19]
Graphene		22–60	0–100 ppm		[50]
Graphene	Devis				
Graphene	Conductance		10–100 ppm	10 s	[50]
LaOCl/SnO <sub>2</sub>	Resistance	350–450	0–4000 ppm		[53]
La/SnO <sub>2</sub>	Resistance	>100	500–2500 ppm		[53]
CNT-polyionic liquids	Resistance		500 ppt		[53,60]
Graphene	Resistance	Room temperature	0–1500 ppm		[61]
Polypyrrole	Resistance	60	100–700 ppm		[62]
1-Ethyl-3-methyl-imidazolium Tetrafluoroborate ([EMIM][BF <sub>4</sub> ]) as ionic liquid and polyvinylidene fluoridehexafluoro propylene (PVDF-HFP) as polymer.	Impedance change	Ambient temperature	0–3000 ppm	50 min	[63]
Graphene oxide	Resistance		0–2000 ppm	3 s	[64]
Pristine graphene/UV light			200 ppt		[65]
Flexible graphene	Resistance	Variable	Low to high		[66]
ZnSb <sub>2</sub> O <sub>6</sub>	Resistance	400 °C	100–400		[68]

### Solid-State Electrochemical Sensor

There have been attempts to develop a solid-state electrochemical sensor for carbon dioxide [71]. Mason et al. [71] developed solid-state reference measuring electrodes and a beta alumina electrolyte substrate with a provision for heating, and made measurements at a temperature of 530 °C. At this temperature, the carbon dioxide sensing was found to be selective from other gases. A novel electrochemical sensor using a ceramic-type component of Na- $\beta$ / $\beta''$  alumina and a reference electrode of glass sealed elemental Na has been developed for measuring CO<sub>2</sub> gas in the temperature range of 300 to 600 °C [72] with a response time of seconds to minutes.

### 3. NO<sub>x</sub> Gas Sensors

Faraday rotation spectroscopy has been proposed for NO<sub>2</sub> determination (Type 4 sensor) [67–73]. It uses a widely tunable external cavity quantum cascade laser (EC-QCL) and operates mode-hop free between 1600 cm<sup>-1</sup> to 1650 cm<sup>-1</sup> and allows Q-branch transition of NO<sub>2</sub> at 1613.2 cm<sup>-1</sup>. A detection limit of 95 ppt has been reported. Meyyappan et al. [73] used a simple casting of single-walled carbon nanotubes on an interdigitated electrode for the detection of NO<sub>2</sub> ranging from sub-ppm to 44 ppb. The response time of the sensor was on the order of seconds.

Among the sensors reported in the literature, there are a few sensors reaching the lowest detection limits of 0.01–0.5 ppb (Table 4). The ZnO [74–76], In<sub>2</sub>O<sub>3</sub> [77], and WO<sub>3</sub> [78,79] sensors showing variable response times fall into this category of low detection limits. A sensor developed with reduced graphene oxide–Cu<sub>2</sub>O nanowire interestingly reached the lowest detection concentration of 0.064 ppm [79–84].

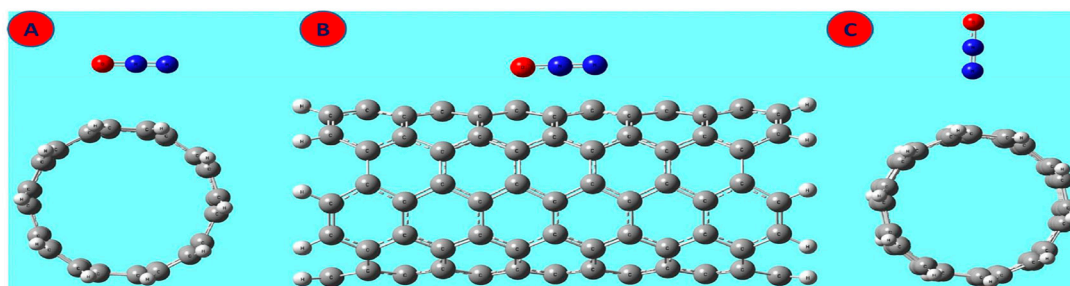
**Table 4.** Several metal–oxide–semiconductor (MOS) materials have been examined for NO<sub>x</sub> detection.

MOS	Operating Temperature (°C)	Response Time (s)	Detection Concentration ppm	Reference
ZnO	RT	240	0.01	[74]
ZnO	200	60	0.1–1	[75]
In <sub>2</sub> O <sub>3</sub>	250	600	0.1–100	[77]
WO <sub>3</sub>	300	80–300	0.5	[78]
WO <sub>3</sub>	300		0.5–2.5	[79]
SnO <sub>2</sub>	300		0.5	[85]
WO <sub>3</sub>	350	180	1	[86]

Nitrous oxide (N<sub>2</sub>O) is released into the atmosphere from chemical plants producing nitric acid and polymers ([https://www.eia.gov/environment/emissions/ghg\\_report/ghg\\_nitrous.cfm](https://www.eia.gov/environment/emissions/ghg_report/ghg_nitrous.cfm)). N<sub>2</sub>O is a colorless toxic pollutant gas with a slightly sweet odor. It is widely used as an anesthetic and analgesic agent in the clinical field, and also as a propellant for pressurized containers in the food industry. It is neither flammable nor explosive. One molecule of N<sub>2</sub>O has the same greenhouse warming power as 300 molecules of carbon dioxide. Two-thirds of anthropogenic N<sub>2</sub>O emissions arise from agricultural soils [86], where N<sub>2</sub>O is formed as part of the bacterial denitrification pathway, in which soil and marine bacteria use oxidized nitrogen compounds as terminal electron acceptors for anaerobic respiration [87]. Once that N<sub>2</sub>O molecule reaches the upper atmosphere, it can stay there for more than 100 years before getting destroyed naturally. Even though nitrous oxide is a moderately undistruptive substance unlike pollutants known as NO<sub>x</sub>, it has recently been reported to participate in the depletion of the ozone layer in the stratosphere ([https://www.eia.gov/environment/emissions/ghg\\_report/ghg\\_nitrous.cfm](https://www.eia.gov/environment/emissions/ghg_report/ghg_nitrous.cfm)). So, it is crucial to control and convert N<sub>2</sub>O to a harmless gas by catalytic surface reactions.

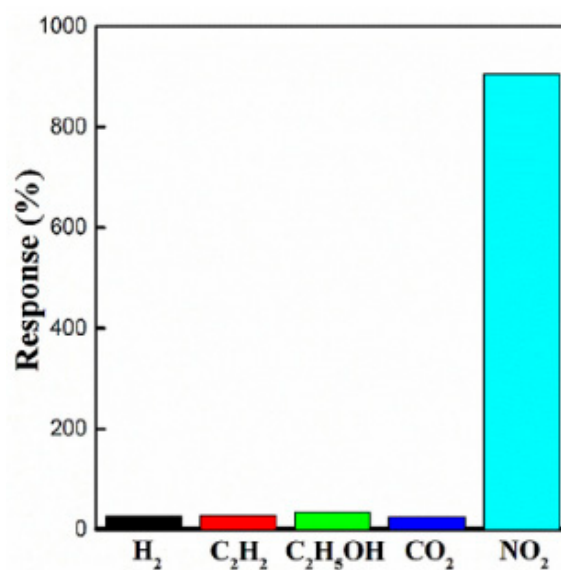
Yoosefian [88] performed density functional studies on the adsorption behavior of nitrous oxide (N<sub>2</sub>O) onto intrinsic carbon nanotube (CNT) and Pd-doped (5,5) single-walled carbon nanotube (Pd-CNT). Pd dopant facilitates the adsorption of N<sub>2</sub>O on the otherwise inert nanotube, as observed from the adsorption energies and global reactivity descriptor values. The adsorption energy of N<sub>2</sub>O on CNT was investigated in three orientations: vertical (VC) and horizontal (HC) to the nanotube

axis, and the nitrogen atom toward the C–C bond (NC). The full optimized structures are indicated in Figure 20.



**Figure 20.** The optimized structures of adsorbed  $\text{N}_2\text{O}$  from three orientations on intrinsic CNT. (A) Vertical (VC) to the nanotube axis; (B) Horizontal (HC) to the nanotube axis; and (C) The nitrogen atom toward the C–C bond (NC). Reproduced with permission from Reference [88].

The adsorption of  $\text{N}_2\text{O}$  changes the electronic conductivity of Pd-CNT, which is attractive for developing the sensor. Recently, heterojunction sensors have been constructed with  $\text{SnO}_2/\text{SnS}_2$ , operating at  $80\text{ }^\circ\text{C}$  (Type 1), sensitive to  $\text{NO}_2$  concentrations in the range of 1–8 ppm [89–96]. The response of the sensor to other analytes is shown in Figure 21.



**Figure 21.** Response of heterojunction sensor to different analytes. Hydrogen ( $\text{H}_2$ ), acetylene ( $\text{C}_2\text{H}_2$ ), ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ), carbon dioxide ( $\text{CO}_2$ ), Nitrogen dioxide ( $\text{NO}_2$ ). Reproduced with permission from reference [89].

Table 5 gives the relative performances of different sensors for  $\text{NO}_2$ .

**Table 5.** The NO<sub>2</sub> sensors fabricated with different materials.

Sensor Material	Operating Temperature	Concentration ppm	Response, S (%)	Response/Recovery Time (s)	Reference
SnO <sub>2</sub> nanofibers	200 °C	1–5	97%/5 ppm	85/110	[94]
SnO <sub>2</sub> core/ZnO shell	RT, UV assisted	1–5	600%/5 ppm	100/220	[95]
Reduced graphene oxide-ZnO	80 °C	1–25	25%/5 ppm	164/499	[96]
Graphene-WO <sub>3</sub>	250 °C	1–20	140%/5 ppm	45/50	[97]
CNT thin film	165 °C	10–100	71%/100 ppb	>30 min	[98]
ZnO nano structure	RT	5–20	1100%/5 ppm	72/69	[99]

In comparison to these sensors, several N<sub>2</sub>O gas sensors have been developed using metal oxides. Table 6 gives the performance of N<sub>2</sub>O sensors (Type 1). These oxide sensors operate at high temperatures in the region of 300–600 °C.

**Table 6.** Metal oxide sensors for N<sub>2</sub>O.

Oxide	Resistance in Air, R <sub>a</sub> (Ω)	Sensitivity to 300 ppm N <sub>2</sub> O	Temperature (°C)
Cr <sub>2</sub> O <sub>3</sub>	3.5 × 10 <sup>5</sup>	0.98	400
Mn <sub>2</sub> O <sub>3</sub>	2.8 × 10 <sup>5</sup>	1.00	400
CO <sub>3</sub> O <sub>4</sub>	1.0 × 10 <sup>4</sup>	0.98	400
NiO	2.8 × 10 <sup>5</sup>	1.00	450
CuO	2.2 × 10 <sup>6</sup>	0.98	300
CdO	5.3 × 10 <sup>9</sup>	1.03	400
MgO	2.4 × 10 <sup>10</sup>	1.00	500
SrO	5.6 × 10 <sup>9</sup>	1.02	400
BaO	2.2 × 10 <sup>10</sup>	1.02	500
In <sub>2</sub> O <sub>3</sub>	3.6 × 10 <sup>4</sup>	1.05	450
SnO <sub>2</sub>	5.8 × 10 <sup>5</sup>	1.66	450
WO <sub>3</sub>	8.5 × 10 <sup>5</sup>	1.32	450
ZnO	2.8 × 10 <sup>6</sup>	1.21	450
TiO <sub>2</sub>	4.9 × 10 <sup>7</sup>	1.01	600
V <sub>2</sub> O <sub>3</sub>	8.7 × 10 <sup>5</sup>	1.00	400
Fe <sub>2</sub> O <sub>3</sub>	3.9 × 10 <sup>7</sup>	1.00	500
GeO <sub>2</sub>	1.2 × 10 <sup>8</sup>	1.00	400
Nb <sub>2</sub> O <sub>5</sub>	9.6 × 10 <sup>8</sup>	1.00	600
MoO <sub>3</sub>	3.1 × 10 <sup>9</sup>	1.00	500
Ta <sub>2</sub> O <sub>5</sub>	2.7 × 10 <sup>10</sup>	1.00	600
La <sub>2</sub> O <sub>3</sub>	1.4 × 10 <sup>9</sup>	1.10	600
CeO <sub>2</sub>	3.9 × 10 <sup>9</sup>	1.01	600
Nd <sub>2</sub> O <sub>3</sub>	5.3 × 10 <sup>9</sup>	1.00	500

Reproduced with permission from Reference [9]; R<sub>g</sub> represents the resistance of the sensor in N<sub>2</sub>O. Sensitivity = (R<sub>g</sub>/R<sub>a</sub>).

#### 4. Methane Sensors

A number of detectors [97–105] have been developed in the last couple of decades for the detection of methane, mainly to detect the leakage of natural gas as the combination of methane and air constitutes an explosive mixture. These detectors are based on optical fiber sensing [98–100], resistive change [98–103] using graphene–polyaniline composite, cataluminescence [104], and refractive index [105,106]. The lower detection limit that can be reached with these methods is about 10 ppm [98–100]. A theoretical calculation of methane adsorption on graphene has been investigated, and it shows an opening of the band structure of graphene upon adsorption [106]. It predicts that the adsorption energy of defected graphene is increased with the number of layers. A low-cost sensor has been developed [107] using a metal oxide that has suitability for detecting methane leaks. The detection of the sensor is in the range of 0.8–2.7 ppm. It is ideally suited for the environmental analysis of methane.

Humayun et al. [108,109] fabricated a highly sensitive, low-cost, and energy-efficient distributed methane (CH<sub>4</sub>) sensor system (DMSS) comprising a metal oxide nanocrystals (MONC)-functionalized MWCNT-based CH<sub>4</sub> chemiresistor sensor (Type 1 sensor). The sensor could sense below 10 ppm of CH<sub>4</sub> in dry air at room temperature. They developed a Gaussian plume triangulation algorithm for the DMSS by which if a geometric model of the surrounding environment is given, then the algorithm can precisely detect and localize a CH<sub>4</sub> leak as well as estimate its mass emission rate. To facilitate faster leak detection, a control algorithm based on the UV-accelerated recovery was developed.

A methane chemiresistive sensor was fabricated using multiwalled carbon nanotubes/metal oxide nano particles that showed a detection limit of 10 ppm. The sensor response to methane and water vapor was measured in this work, showing that the sensor was inactive to water. A modulation spectroscopy sensor (Type 4) using a multi-pass cell with a detection sensitivity of 1.2 ppb has recently been reported for methane [109,110]. Using this sensor, the concentration of methane in ambient air was determined to be about 2.01 ppm. With this concentration of methane in the atmosphere referring to Figure 3, the percentage error in CO<sub>2</sub> measurement will be about 4% unless the CO<sub>2</sub> selected sensor is not influenced by methane gas.

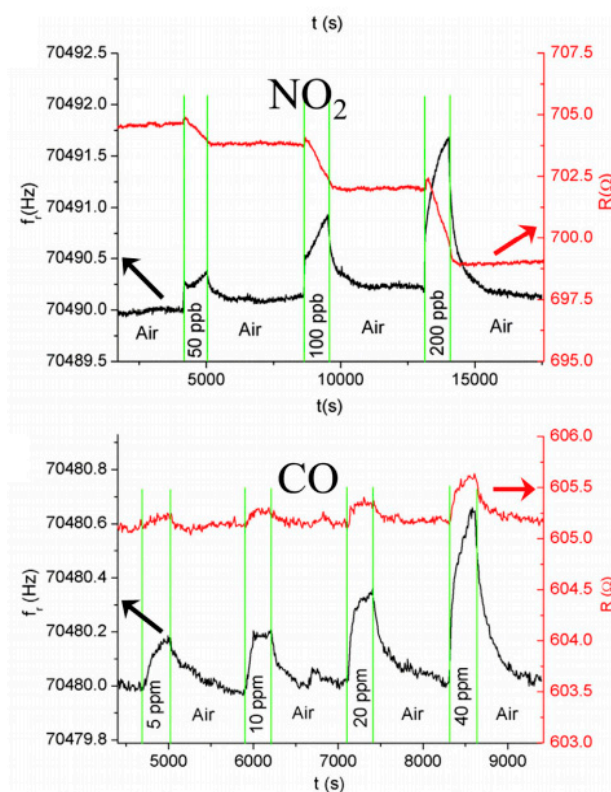
### 5. Modern Technology Using 3D and Ink Jet Printing of Gas Sensors

With rapid developments in finding active materials suitable for greenhouse gas sensors, there is need for consideration of the cost of making these devices. In this context, the 3D-printing of the active material would be of interest. High-quality 3D-printed desk-top devices have been produced at a low cost without conventional micromechanical systems. Taylor and Velásquez-García [111] have reported a novel electrospray printed nanostructured graphene oxide for gas detection. A number of gases were examined in this work with a variety of detection limits. For CO<sub>2</sub>, the detection limit was set at 1000 ppm. The preliminary reports opened up opportunities to modify the spray printing for extension to lower detection limits. Rieu et al. [112] developed an ink jet-printed SnO<sub>2</sub> gas sensor on a plastic substrate. Both the gas-sensitive layer and the heating transducer were ink jet-printed. This method of making the sensor compares well with sol-gel tin oxide film. The following detection limits have been reported for CO and NO<sub>2</sub> (Table 7) The limit of detection fell far above the other methods discussed in the earlier sections. Furthermore, the operating temperature of the sensor was in the range of 200–300 °C (Table 7). Sulfonated graphene was used for detecting NO<sub>2</sub> at room temperature by Liua et al. [113]. Sulfonated graphene and SnO<sub>2</sub> particles are combined to form the active material of the sensor. The process involves the direct deposition of SnO<sub>2</sub> nanoparticles on reduced graphene oxide. A high performance and a good sensitivity were achieved by this active material. A detection limit of 1 ppm of NO<sub>2</sub> was reached with this sensor. The room-temperature detection using sulfonated graphene opens up opportunities for developing other greenhouse gas sensors. A screen-printed piezoelectric microcantilever (Type 6 sensor) has been used to detect and determine NO<sub>2</sub> and CO gases using oxygen plasma-treated multiwalled carbon nanotubes as the sensitive layer [114]. Figure 22 shows the response of the cantilever for the two gases, with both positive and negative shifts in resonant frequency, attributed to stress and mass effects at different analyte concentrations.

**Table 7.** Limits of detection of CO and NO<sub>2</sub> using the SnO<sub>2</sub> ink jet-printed sensors.

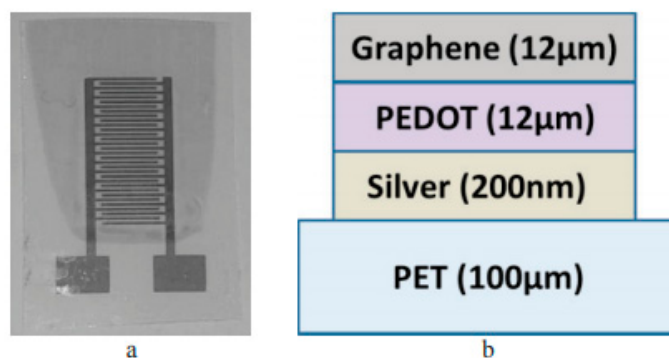
Gas	Detection Level (ppm)			
CO (ppm)	0.5 (300 °C, dry air)	24 (300 °C, wet air)	0.4 (200 °C, dry air)	46 (200 °C, wet air)
NO <sub>2</sub> (ppm)	6 (300 °C, dry air)	9 (300 °C, wet air)	1 (200 °C, dry air)	6 (200 °C, wet air)





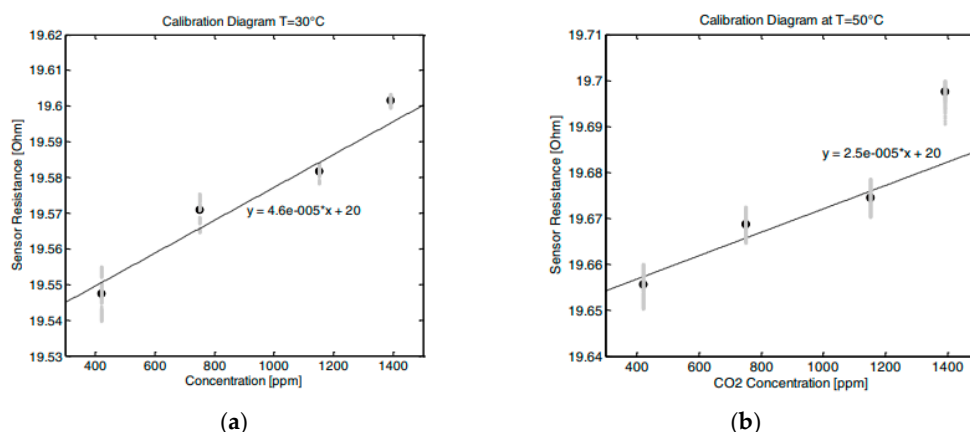
**Figure 22.** Screen-printed cantilever with multiwalled carbon nanotube functionalized by oxygen-argon plasma as sensitive material. Reproduced with permission from Reference [114].

The sensor's response to  $\text{NO}_2$  was remarkable at ppb levels. A new generation of sensors for greenhouse gases utilizing 3D printing and ink jet printing is to be expected in the future, and are being researched in the literature [114–119]. Although some of the analytes described in this section do not fall into the class of greenhouse gases, they provide the pathway for the development of new greenhouse gas sensors through ink jet printing. With this technology, the sensors will have high reproducibility, sensitivity, and low cost. A 3D-printed  $\text{CO}_2$  sensor was constructed with a double layer of PEDOT/PSS (poly(3,4-ethylenedioxythiophene)/polystyrene sulfonate) and graphene on interdigitated electrodes that was made by inkjet printing using conductive silver nanoparticles. The sensor works on the principle illustrated for type 1 and has sensitivity of  $45 \mu\text{ohm/ppm}$  [117]. Figure 23 shows the different layers in the sensor and the interdigitated fingers.



**Figure 23.** (a) Device dimensions: 1.8 cm by 1 cm with a thickness of about  $140 \mu\text{m}$ . IDT fingers: width ( $200 \mu\text{m}$ ), length (6 mm), thickness (200 nm), track spacing ( $350 \mu\text{m}$ ); (b) Schematic of the sensor structure PET: Polyethylene terephthalate. Reproduced with permission from Reference [117].

The sensor's responses at two different temperatures are given in Figure 24.

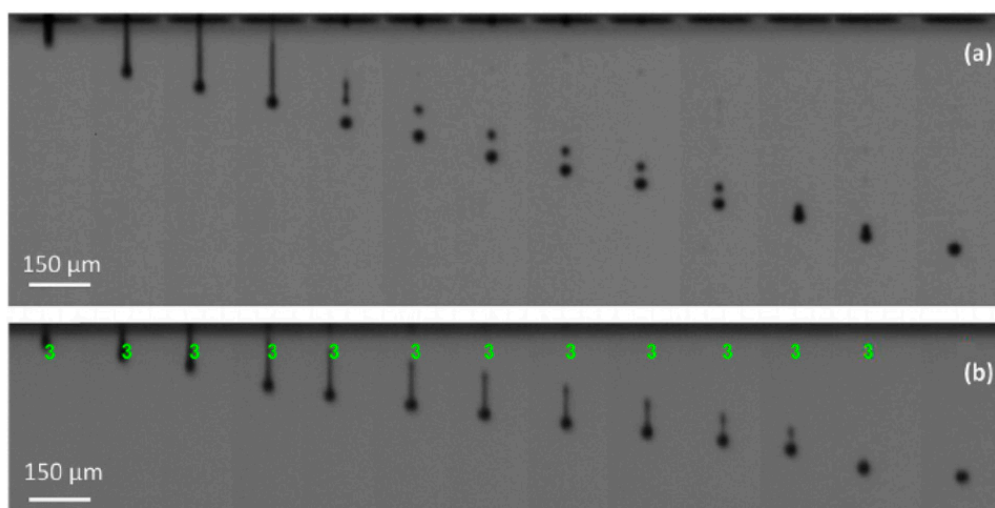


**Figure 24.** An inkjet printed sensor response to CO<sub>2</sub> gas at two different temperatures. Reproduced with permission from reference [117].

3D inkjet printing has provided an economical manufacturing technique for the CO<sub>2</sub> sensor [118]. It uses a colloidal dispersion in the printing ink and is composed of La<sub>0.8</sub>Sr<sub>0.2</sub>MnO<sub>3</sub> (LSM)(Y<sub>2</sub>O<sub>3</sub>)<sub>0.08</sub>(ZrO<sub>2</sub>)<sub>0.92</sub>(YSZ). It uses an electrochemical cell Ni-YSZ | YSZ | YSZ-LSM | LSM, with YSZ (yttria stabilized zirconia) as electrolyte sintered to a Ni-yttria stabilized zirconia (Ni-YSZ). The carbon dioxide reaction in the electrochemical fuel cell is



The inkjet 3D printing performance is dependent on the formation and ejection of droplets. Its propagation depends on acoustic pressure waves in the fluid held in the chamber behind the printing nozzle and in the YSZ-LSM LSM ((La<sub>1-x</sub>Sr<sub>x</sub>)<sub>1-y</sub>MnO<sub>3</sub>) sensor. The droplets time interval is shown in Figure 25. The colloidal LSM dispersions are stabilized using electrostatic dispersants which are commercially available from different manufacturers, as shown in Table 8.

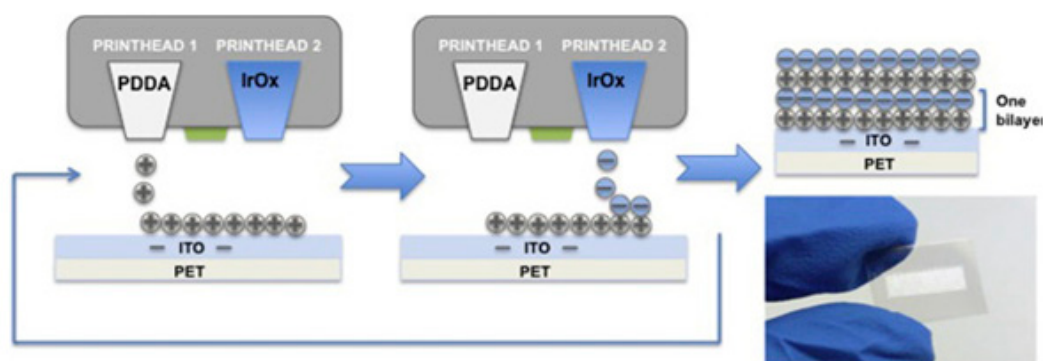


**Figure 25.** (a) YSZ (Yttria-stabilized Zirconia) ink and (b) Butyl alcohol-based YSZ-LSM ((La<sub>1-x</sub>Sr<sub>x</sub>)<sub>1-y</sub>MnO<sub>3</sub>) ink droplets from a single nozzle, with time intervals of 10 μs and 5 μs, respectively. The nozzle number ("3") of the printhead is shown in (b). Reproduced with permission from Reference [118].

**Table 8.** Electrostatic Dispersants for Inkjet printing.

Dispersant	Trade Name	Supplier
Aurintricarboxylic acid ammonium salt	Aluminon	Sigma-Aldrich, Saint Louis, MI, USA
Ammonium citrate		Sigma-Aldrich, Saint Louis, MI, USA
Sodium sulfonate	Trion	Sigma-Aldrich, Saint Louis, MI, USA
Ammonium polyacrylate	Dispex-A40	BASF, Ludwigshafen, Germany
Polymethylmethacrylate	Darvan-C	Vanderbilt Minerals, Saint Lawrence County, NY, USA
Ammonium polyacrylate	Dispex-A40	BASF, Ludwigshafen, Germany

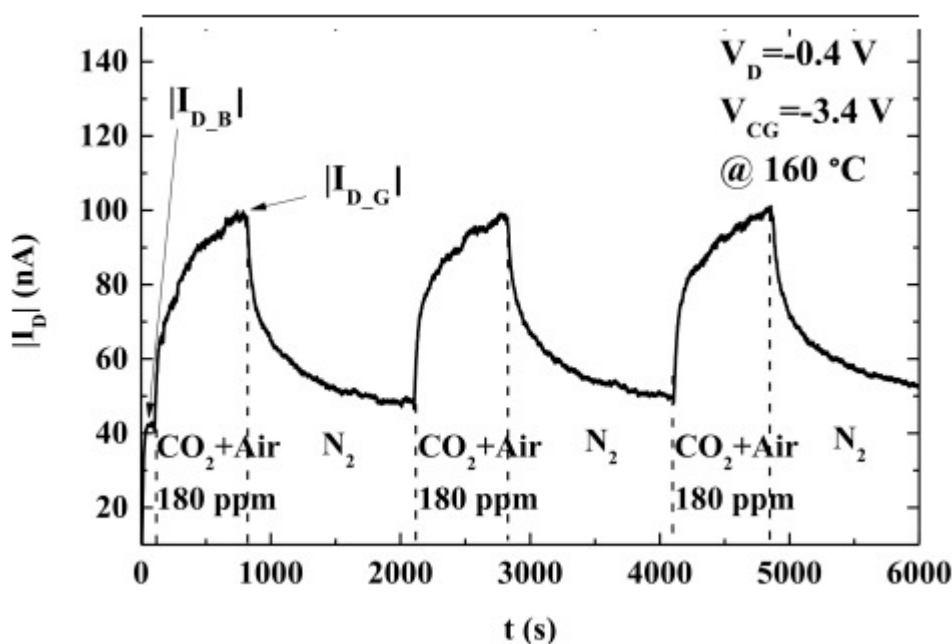
The performance of the fuel cell is linked to the concentration of CO<sub>2</sub>, and it is in the initial stages of development. As discussed in the previous section on carbon dioxide, a pH electrode can be used as a CO<sub>2</sub> sensor. A large-scale layer-by-layer ink jet printing of flexible iridium oxide for a hydrogen ion sensor (Figure 26) has been reported [119]. The sensor's active material is iridium oxide, which interacts with the analyte.



**Figure 26.** Deposition of positively charged poly(diallyldimethylammonium chloride) (PDDA, 20% wt. in water, medium molecular weight MW 200,000–350,000) and positively charged particles by inkjet printing. ITO: indium tin oxide glass; IrOx: iridium oxide; PET: polyethylene terephthalate. The feedback arrow indicates layer-by-layer deposition. Reproduced with permission from Reference [119].

A flexible tin oxide gas sensor was developed by an inkjet printing process. Here the active material was tin oxide, printed on gold deposited in a digitated electrode spacing of 500 and 200  $\mu\text{m}$ . Carbon monoxide was the analyte, and the measurements were done at 300  $^{\circ}\text{C}$  in the concentration range of 20–200 ppm [120].

A type 3 sensor for CO<sub>2</sub> was recently been reported [121] using a sodium ion conducting solid electrolyte having a floating gate passivated by an insulator stack and a control gate made of an interdigitated field gate (Figure 27). The sensing material, (3-aminopropyl) triethoxysilane (APTES), was added by inkjet printing. The sensor operated at 160  $^{\circ}\text{C}$  in the concentration range of 125–325 ppm.



**Figure 27.** Type 3 inkjet-printed sensor for CO<sub>2</sub> operating at 160 °C. I<sub>D</sub>: drain current. Responds to CO<sub>2</sub> concentration. V<sub>D</sub> and V<sub>CG</sub> are drain and gate voltages. Reproduced with permission from Reference [121].

Inkjet printing technology has very high commercial value, and the developments have been rapidly pursued for long-term testing. Inkjet-printed graphene sensors (Type 1) for nitrogen oxide gas and inkjet printed platinum decorated tungsten oxide nanoparticles on silicon substrate for gas sensing (Type 1) have been reported [122,123]. The methodologies reported in these publications are interesting, and may be extendable to greenhouse gases in the future [124–126].

## 6. Outlook and Conclusions

The recent reports of monitoring greenhouse effect gases have conclusively shown that CO<sub>2</sub> levels in the atmosphere have been steadily increasing. The UN reports have consistently been promoting the use of renewable energies for controlling the atmospheric conditions. The atmospheric conditions need to be steadily monitored, and for this purpose we need to have reliable measurements of the greenhouse effect gases. The acquisition of the data requires the development of cheap, robust, and accurate sensors. Although the area of sensors has been growing rapidly, sensors directed towards the greenhouse effect gases using nano materials have been recent. This survey of the greenhouse effect gas sensors suggests that there has been active interest in developing gas sensing technologies. The first part of the paper projects the progressive increase in the concentrations of greenhouse gases in the atmosphere and its consequences. The second part discusses the developments in the areas of carbon dioxide, nitrogen oxide, and methane sensors using carbon nanotubes and polymer composites. A few reports on using graphene, which is a material having large surface area relative to carbon nanotubes, with high electrical and thermal conductivities, opens opportunities for developing novel greenhouse gas sensors. The current trend is to use 3D printing or ink jet printing in making the sensor devices. This survey of greenhouse gas sensors shows that so far the developments have been based on discovering new active materials and examining their characteristics limited to either pure gases with limited interferences. There are very few sensors that would respond to only one greenhouse gas. This short review is to focus on the need to have reliable sensors for greenhouse gases in every city or town for a periodical review of their concentrations.

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